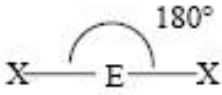
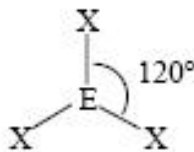
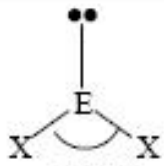
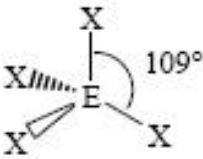
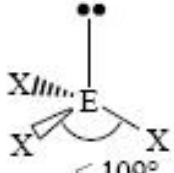

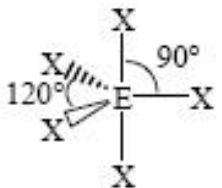
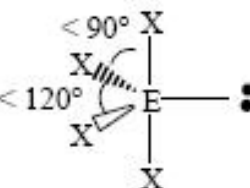
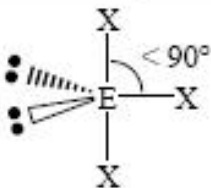
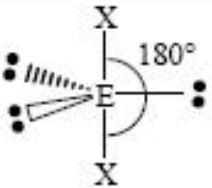
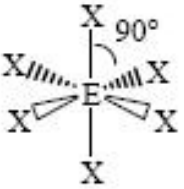
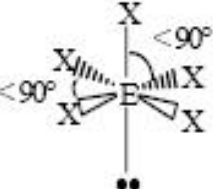
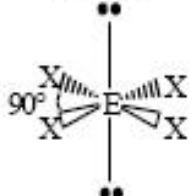
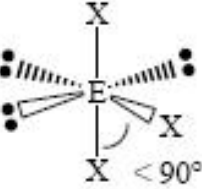
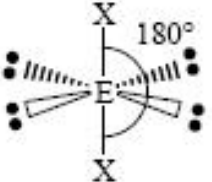


VSEPR Geometries

Steric No.	Basic Geometry 0 lone pair	1 lone pair	2 lone pairs	3 lone pairs	4 lone pairs
2	 <p style="text-align: center;">Linear</p>				
3	 <p style="text-align: center;">Trigonal Planar</p>	 <p style="text-align: center;">Bent or Angular</p>			
4	 <p style="text-align: center;">Tetrahedral</p>	 <p style="text-align: center;">Trigonal Pyramid</p>	 <p style="text-align: center;">Bent or Angular</p>		
5	 <p style="text-align: center;">Trigonal Bipyramid</p>	 <p style="text-align: center;">Sawhorse or Seesaw</p>	 <p style="text-align: center;">T-shape</p>	 <p style="text-align: center;">Linear</p>	
6	 <p style="text-align: center;">Octahedral</p>	 <p style="text-align: center;">Square Pyramid</p>	 <p style="text-align: center;">Square Planar</p>	 <p style="text-align: center;">T-shape</p>	 <p style="text-align: center;">Linear</p>

ELECTRON CHARGE 5 CENTER	Bonding Pair	Lone Pair	PCl_5 	Geometry 	5 bond pair ↓ Bond angle - 90°, 120° ↓ Trigonal Bipyramidal ✓
	4	1	SF_4 TeCl_4 $(\text{IF}_4)^+$ XeO_2F_2 	Geometry 	4 bond pair 1 lone pair ↓ Bond angle < 90°, < 120° ↓ Seesaw ✓
	3	2	ClF_3 ICl_3 BrF_3 $(\text{XeF}_3)^+$ 	Geometry 	3 bond pair 2 lone pair ↓ Bond angle < 90° ↓ T shape ✓
	2	3	$(\text{I}_3)^-$ $(\text{ICl}_2)^-$ XeF_2 	Geometry 	2 bond pair 3 lone pair ↓ Bond angle 180° Linear ✓